The reason for the loss of activity of titanium anodes coated with a layer of \( \text{RuO}_2 \) and \( \text{TiO}_2 \)

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The galvanostatic method has been employed for following the time change of the ATA potential in 0.5 N \( \text{H}_2\text{SO}_4 \). On the basis of the experimental data it is proposed that the loss of ATA activity results from the formation of a non-conductive oxide film at the titanium-active layer interface.

1. Introduction

During the last years there have been many studies of the electrochemical properties of titanium anodes coated with a layer of ruthenium and titanium oxides, i.e. activated titanium anodes (ATA). With respect to the extensive industrial use of these anodes it is important to know why they lose their electrochemical activity after a certain time. (This is manifested by increasing anode potentials). This knowledge could then be exploited in the course of producing ATA, or possibly during the production of electrodes from used anodes.

According to certain data [1] the loss of electrochemical activity of the ATA occurs because of the oxidation of ruthenium dioxide to ruthenium tetroxide and thus, the dissolution of one component of the active layer. According to thermodynamical calculations [2], the oxidation of \( \text{RuO}_2 \) to \( \text{RuO}_4 \) occurs at potentials more positive than 1.387 V (versus the standard hydrogen electrode). According to other data [3, 4], it seems possible that in the course of electrolysis a layer of the titanium oxide having a valve effect is formed. This layer would also increase the anode potential.

In the present work the reasons for the loss of the ATA activity were studied in detail.

2. Experimental

A galvanostatic method (\( i = \text{constant} \)) was used for ATA activity measurements. A SVÚOM–VPZ potentiostat in the galvanostatic connection served as a source of the constant current. The potential difference between the working and reference electrode was recorded on a Rikadenki (Kogyo Co. Ltd) six-line recorder. The same instrument was used to record the current between the working and the auxiliary electrode. The change of the potential with time served as a measure of the ATA activity.

All the measurements were performed at constant temperature, in an ordinary glass vessel with a three-electrode connection. The mercurous sulphate electrode in 0.5 N \( \text{H}_2\text{SO}_4 \) was chosen as the reference electrode. A cylindrical Pt gauze with a geometrical surface of about 30 cm\(^2\) served as the auxiliary electrode. An activated Ti wire, 2 mm in diameter, served as the working electrode. The activated titanium wire was situated at the centre of the cylindrical auxiliary electrode, in order to obtain an even distribution of the anodic current density.

Two types of activated titanium anode were prepared. These were anodes with a layer of the \( \text{RuO}_2 \) only and anodes with a layer of \( \text{RuO}_2 \) and \( \text{TiO}_2 \), the ratio of the metals being 1:1 by mass. The electrodes were prepared by applying ethanol solutions of \( \text{RuCl}_3 \) or \( \text{RuCl}_3 + \text{TiCl}_3 \), respectively, on to a titanium wire, which was subsequently heated at a temperature above 400° C.

All the measurements were made in 0.5 N \( \text{H}_2\text{SO}_4 \) solution prepared from analytical grade sulphuric acid and distilled water. During each measurement the temperature was 20 ± 1° C. The measurement consisted in following the time change of the working electrode potential for
different current densities. The current densities in the range 0.1–1.5 A cm\(^{-2}\) were used.

3. Results

A typical relationship between the potential of an activated titanium anode in 0.5 N H\(_2\)SO\(_4\) and time for a current density of 0.5 A cm\(^{-2}\) is shown in Fig. 1. In this case the active layer on the titanium contains both RuO\(_2\) and TiO\(_2\) (metal ratio 1 : 1).

![Fig. 1. The time change of the ATA potential in 0.5 N H\(_2\)SO\(_4\) at 0.5 A cm\(^{-2}\) (RuO\(_2\) + TiO\(_2\)).](image)

The change of the activated titanium anode potential is defined by the Equation

\[
\Delta \phi = \phi - \phi_0
\]

where \(\phi_0\) is a potential of the ATA at time \(t = 0\), \(\phi\) is a potential of the ATA at time \(t\). The ATA potential remains unaltered for quite a long time but then increases rapidly. At first sight, these curves are similar to those corresponding to the potential changes on electrodes forming non-conducting oxidic films on their surfaces [5].

![Fig. 2. Dependence of \(\Delta \phi\) on \(\Delta \phi t\) for an ATA (RuO\(_2\) + TiO\(_2\)) in 0.5 N H\(_2\)SO\(_4\) for a current density of 0.5 A cm\(^{-2}\).](image)

Curves of the same type are obtained for all current densities in the range 0.1–1.5 A cm\(^{-2}\) for the ATA with a layer formed by RuO\(_2\) and TiO\(_2\), or with RuO\(_2\) itself, albeit with different values of the constants \(a\) and \(k\).

By plotting the current density \(i\) as a function of \(\ln k\) we can obtain the straight-line relationship shown in Fig. 3, for which

\[
i = a_0 + b \ln k
\]

or

\[
k = a_i \exp \frac{1}{b}
\]

where \(b\) is the slope (Fig. 3). For the ATA with a layer of RuO\(_2\) and TiO\(_2\) \(b = 0.311\) A cm\(^{-2}\).

In the industrial use of ATA, the service time* represents one of the governing factors. In the present work the service time is defined as a time when, on a theoretical basis, the potential increases to an infinite value. This value can also be determined experimentally with a relatively high precision. The service time determined in this way is denoted by \(\tau\). A plot of service time versus current density is shown in Fig. 4. As can be expected, the service time drops with increasing current density.

The service time of the ATA can also be defined by means of the charge \(Q\) which passes through the electrode for the time \(\tau\). Then

\[
Q = ir
\]

where \(Q\) is the time of operation before the anode potential increases significantly (i.e. by 200–300 mV).